

Test Bank for Anatomy & Physiology The Unity of Form and Function 10th Saladin

TEST BANK SAMPLE PREVIEW

PUBLISHER

McGraw-Hill

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RESOURCE TYPE

Test Bank

ISBN

9781265328627

CORRECT ANSWERS ARE LOCATED IN THE 2ND HALF OF THIS DOC.

TRUE/FALSE - Write 'T' if the statement is true and 'F' if the statement is false.

- 1) Minerals are organic elements extracted from the soil by plants.
 - ☐ true
 - ☐ false
- 2) Molecules composed of two or more atoms are called compounds.
 - ☐ true
 - ☐ false
- 3) Hydrogen, deuterium, and tritium are three isotopes of hydrogen.
 - ☐ true
 - ☐ false
- 4) Potassium, sodium, and chlorine are trace elements.
 - ☐ true
 - ☐ false
- 5) Ionic bonds break apart in water more easily than covalent bonds do.
 - ☐ true
 - ☐ false
- 6) A solution is a mixture of two or more substances that are physically blended but *not* chemically combined.
 - ☐ true
 - ☐ false
- 7) The pH of blood plasma is approximately 7.4, which is slightly acidic.
 - ☐ true
 - ☐ false
- 8) The high heat capacity of water makes it a very ineffective coolant.
 - ☐ true
 - ☐ false
- 9) In an exchange reaction, covalent bonds are broken and new covalent bonds are formed.
 - ☐ true
 - ☐ false

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- 10) Chemical reactions in which larger molecules are broken down into smaller ones are called catabolic reactions.
- ☐ true
 - ☐ false
- 11) The opposite of a dehydration synthesis reaction is a hydrolysis reaction.
- ☐ true
 - ☐ false
- 12) Unsaturated fatty acids have as much hydrogen as they can carry.
- ☐ true
 - ☐ false
- 13) A dipeptide is a molecule with two peptide bonds.
- ☐ true
 - ☐ false
- 14) All amino acids have both a carboxyl group and an amino group attached to a central carbon.
- ☐ true
 - ☐ false
- TBEXAM.COM
- 15) ATP is the body's most important form of long-term energy storage.
- ☐ true
 - ☐ false
- 16) A molecule that is oxidized gains electrons and energy.
- ☐ true
 - ☐ false
- 17) Minerals are organic molecules that must be obtained through food.
- ☐ true
 - ☐ false

CHECK ALL THE APPLY. Choose all options that best completes the statement or answers the question.

18) Which of these is a cation?

- A) O_2
- B) K^+
- C) Na^+
- D) Ca^{2+}
- E) Cl^-

MULTIPLE CHOICE - Choose the one alternative that best completes the statement or answers the question.

19) The most abundant element in the human body, by weight, is_____.

- A) nitrogen
- B) hydrogen
- C) carbon
- D) oxygen
- E) calcium

20) Sodium has an atomic number of 11 and an atomic mass of 23. Sodium has_____.

- A) 12 neutrons and 11 protons
- B) 12 protons and 11 neutrons TBEXAM.COM
- C) 12 electrons and 11 neutrons
- D) 12 protons and 11 electrons
- E) 12 electrons and 11 protons

21) The chemical properties of an atom are determined by its_____.

- A) protons
- B) electrons
- C) neutrons
- D) protons and neutrons
- E) particles

22) Na (atomic no. 11) reacts with Cl (atomic no. 17) to become stable. In the reaction, Na will_____, while Cl will_____.

- A) accept one electron; give up one electron
- B) give up one proton; accept one proton
- C) share one electron with chlorine; share one electron with sodium
- D) become an anion; become a cation
- E) give up one electron; accept one electron

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- 23) Oxygen has an atomic number of 8 and an atomic mass of 16. How many valence electrons does it have?
- A) 2
 - B) 4
 - C) 6
 - D) 8
 - E) 16
- 24) Oxygen has an atomic number of eight. When two oxygen atoms come together, they form a(n) _____ bond.
- A) hydrogen
 - B) nonpolar covalent
 - C) polar covalent
 - D) ionic
 - E) Van der Waals
- 25) When table salt, sodium chloride (NaCl), is placed in water _____.
- A) Na^+ and Cl^- form ionic bonds with each other
 - B) Na^+ and Cl^- form polar covalent bonds with each other
 - C) Na^+ and Cl^- form hydrogen bonds with water
 - D) ionic bonds between Na^+ and Cl^- are broken
 - E) Na^+ and Cl^- become separated by their Van der Waals forces
- 26) The bonding properties of an atom are determined by its _____.
- A) electrons
 - B) protons
 - C) positrons
 - D) neutrons
 - E) photons
- 27) What type of bond attracts one water molecule to another?
- A) An ionic bond
 - B) A peptide bond
 - C) A hydrogen bond
 - D) A covalent bond
 - E) A hydrolytic bond

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- 28) _____ account for 98.5% of the body's weight.
- A) Carbon, oxygen, hydrogen, sodium, potassium, and chlorine
 - B) Carbon, oxygen, iron, sodium, potassium, and chlorine
 - C) Carbon, nitrogen, hydrogen, sodium, potassium, and chlorine
 - D) Carbon, oxygen, hydrogen, nitrogen, sodium, and potassium
 - E) Carbon, oxygen, hydrogen, nitrogen, calcium, and phosphorus
- 29) _____ differ from one another in their number of neutrons and atomic mass.
- A) Cations
 - B) Anions
 - C) Isotopes
 - D) Electrolytes
 - E) Free radicals
- 30) When jumping into water you notice resistance. This resistance is caused by water's_____.
- A) adhesiveness
 - B) cohesiveness
 - C) hydrophobic tension
 - D) hydrophilic tension
 - E) osmotic equilibrium
- 31) Which of these is hydrophobic?
- A) Glucose
 - B) K^+
 - C) Cl^-
 - D) Water
 - E) Fat
- 32) Blood contains NaCl, protein, and cells. The NaCl is in a(n)_____, the protein is in a(n)_____, and the cells are in a_____.
- A) emulsion; solution; suspension
 - B) solvent; emulsion; colloid
 - C) colloid; suspension; solution
 - D) suspension; colloid; solution
 - E) solution; colloid; suspension

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33) Which of these is the most appropriate to express the number of molecules per volume?

- A) Molarity
- B) Volume
- C) Percentage
- D) Weight per volume
- E) Milliequivalents per liter

34) A solution with pH 4 has _____ the H^+ concentration of a solution with pH 8.

- A) $\frac{1}{2}$
- B) 2 times
- C) 4 times
- D) 10,000 times
- E) 1/10,000

35) Which of these has the highest H^+ concentration?

- A) Lemon juice, pH = 2.3
- B) Red wine, pH = 3.2
- C) Tomato juice, pH = 4.7
- D) Saliva, pH = 6.6
- E) Household ammonia, pH = 10.8

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36) In a workout your muscle cells produce lactate, yet you maintain a constant blood pH because _____.

- A) metabolic acids are neutralized in muscle cells before released into the blood
- B) metabolic bases are produced at the same rate by muscle cells to neutralize the acids
- C) the respiratory system removes excess H^+ from the blood before the pH is lowered
- D) the body contains chemicals called buffers that resist changes in pH
- E) endothelial cells secrete excess H^+ to prevent a decrease in pH

37) A solution that resists a change in pH when an acid or base is added to it is a(n) _____.

- A) buffer
- B) catalyst
- C) reducing agent
- D) oxidizing agent
- E) colloid

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- 38) A chemical reaction that removes electrons from an atom is called a(n)_____ reaction.
- A) reduction
 - B) condensation
 - C) hydrolysis
 - D) anabolic
 - E) oxidation
- 39) The breakdown of glycogen (an energy-storage compound) is an example of a(n)_____ reaction.
- A) exergonic
 - B) endergonic
 - C) exchange
 - D) synthesis
 - E) equilibrium
- 40) The most relevant free energy in human physiology is the energy stored in_____.
- A) electrolytes ionized in water
 - B) free radicals with an odd number of electrons
 - C) radioisotopes
 - D) the chemical bonds of organic molecules
 - E) Van der Waals forces
- 41) Potential energy stored in bonds is released as_____ energy.
- A) electromagnetic
 - B) electrical
 - C) chemical
 - D) heat
 - E) kinetic
- 42) The breakdown of glucose to yield carbon dioxide, oxygen, and ATP can be described as_____.
- A) anabolic and endergonic
 - B) catabolic and exergonic
 - C) anabolic and exergonic
 - D) catabolic and endergonic
 - E) anabolic and exothermic

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- 43) Which one of the following would *not* increase the rate of a reaction?
- A) An increase in reactant concentrations
 - B) A rise in temperature
 - C) The presence of a catalyst
 - D) The presence of an enzyme
 - E) A decrease in reactant concentrations
- 44) Which of the following terms encompasses all of the other ones?
- A) Catabolism
 - B) Anabolism
 - C) Metabolism
 - D) Oxidation reactions
 - E) Reduction reactions
- 45) The breakdown of starch by digestive enzymes into glucose molecules is a(n)_____ reaction.
- A) synthesis
 - B) decomposition
 - C) exchange
 - D) anabolic
 - E) reduction
- 46) Which of the following equations depicts an exchange reaction?
- A) $AB \rightarrow A + B$
 - B) $A + B \rightarrow AB$
 - C) $AB + CD \rightarrow AC + BD$
 - D) $AB \rightarrow A^- + B^+$
 - E) $A + B \rightarrow AB \rightarrow C + D$
- 47) Which of these functional groups contains nitrogen?
- A) Carboxyl group
 - B) Methyl group
 - C) Hydroxyl group
 - D) Amino group
 - E) Phosphate group

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48) Which of the following is *not* an organic compound?

- A) $C_{16}H_{18}N_3ClS$
- B) $Na_2HPO_3(H_2O)_5$
- C) CH_4
- D) $C_3H_7O_2N$

49) A _____ reaction breaks a _____ down into its monomers.

- A) hydrolysis; polymer
- B) dehydration synthesis; molecule
- C) dehydration synthesis; polymer
- D) polymer; molecule
- E) condensation; reactant

50) The formula of an amino group is _____; the formula of a carboxyl group is _____.

- A) $-COOH$; $-OH$
- B) $-CH_3$; $-NH_2$
- C) $-OH$; $-SH$
- D) $-NH_2$; $-COOH$
- E) $-SH$; $-H_2PO_4$

51) Table sugar is a disaccharide called TBEXAM.COM and is made up of the monomer(s) _____.

- A) maltose; glucose and sucrose
- B) sucrose; glucose and fructose
- C) lactose; glucose and galactose
- D) glycogen; glucose and fructose
- E) glucose; galactose and fructose

52) Which of the following is a disaccharide?

- A) Galactose
- B) Lactose
- C) Glucose
- D) Fructose
- E) Amylose

53) _____ is a monosaccharide, whereas _____ is a polysaccharide.

- A) Fructose; sucrose
- B) Galactose; maltose
- C) Lactose; glycogen
- D) Glucose; starch
- E) Cellulose; glucose

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- 54) In general, _____ have a 2:1 ratio of hydrogen to oxygen.
- A) enzymes
 - B) proteins
 - C) lipids
 - D) carbohydrates
 - E) nucleic acids
- 55) Proteoglycans are composed of _____.
- A) carbohydrates and fats
 - B) nucleic acids and fats
 - C) carbohydrates and proteins
 - D) proteins and fats
 - E) nucleic acids and proteins
- 56) Triglycerides consist of a 3-carbon compound called _____ bound to three _____.
- A) pyruvate; fatty acids
 - B) lactate; glycerols
 - C) eicosanoid; steroids
 - D) glycerol; fatty acids
 - E) sterol; fatty acids
- TBEXAM.COM
- 57) _____ are major components of cell membranes, and are said to be _____.
- A) Triglycerides; hydrophobic
 - B) Steroids; hydrophilic
 - C) Bile acids; fat-soluble
 - D) Eicosanoids; water-soluble
 - E) Phospholipids; amphiphilic
- 58) Which of these molecules is hydrophobic?
- A) Glucose
 - B) Cholesterol
 - C) Amino acid
 - D) Protein
 - E) Disaccharide

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- 59) Proteins perform all of the following functions *except* _____.
A) catalyze metabolic reactions
B) give structural strength to cells and tissues
C) produce muscular and other forms of movement
D) regulate transport of solutes into and out of cells
E) store hereditary information
- 60) A drastic conformational change in a protein in response to extreme heat or pH is called _____.
A) contamination
B) denaturation
C) saturation
D) sedimentation
E) deconformation
- 61) Proteins are _____ built from _____ different amino acids. 01_20_2015_CS-3282
A) monomers; 10
B) molecules; 10
C) polymers; 20
D) macromolecules; 40
E) peptides; 25
- 62) The folding and coiling of a protein into a globular shape is the _____ structure of the protein.
A) primary
B) secondary
C) tertiary
D) quaternary
E) denatured
- 63) An enzyme is substrate-specific because of the shape of its _____.
A) active site
B) receptor
C) secondary structure
D) terminal amino acid
E) alpha chain

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64) Lactose is the substrate of which enzyme?

- A) Lactase
- B) Amylase
- C) Galactase
- D) Protease
- E) Sucrase

65) All enzymes are_____.

- A) cofactors
- B) proteins
- C) lipids
- D) carbohydrates
- E) nucleic acids

66) Nucleic acids are_____ of_____.

- A) monomers; monosaccharides
- B) monomers; ATP
- C) polymers; nucleotides
- D) polymers; cAMP
- E) polymers; DNA

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67) ATP_____ endergonic and exergonic reactions.

- A) opposes
- B) decomposes
- C) reduces
- D) links
- E) dehydrates

68) An atom with 12 electrons, 13 neutrons, and 11 protons is a(n) _____.

- A) anion
- B) cation
- C) free radical
- D) both an anion and a free radical
- E) both a cation and a free radical

69) The concentration of a solution may be expressed by all of the following *except*_____.

- A) weight per volume
- B) percentage
- C) molarity
- D) pH

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70) The vibration of an ear drum is an example of _____ energy.

- A) kinetic
- B) potential
- C) elastic
- D) radiant

71) In the following reaction, what is(are) the product(s)? $\text{CO}_2 + \text{H}_2\text{O} \rightarrow \text{H}_2\text{CO}_3$

- A) H_2CO_3
- B) CO_2 and H_2O
- C) CO_2 and H_2CO_3
- D) H_2O and H_2CO_3

72) Which of the following will increase the rate of a chemical reaction?

- A) An increase in reactant concentration
- B) An increase in product concentration
- C) A decreased temperature
- D) Enzyme inhibition

73) Carbon is very versatile in forming bonds with other atoms because it has _____ valence electrons.

- A) four
- B) two
- C) eight
- D) six

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74) Amylase is a digestive enzyme that breaks starches down into sugars through _____ reactions.

- A) hydrolysis
- B) dehydration synthesis
- C) anabolic
- D) endergonic

75) Which of the following is **not** a nucleotide?

- A) RNA
- B) GTP
- C) ATP
- D) cAMP

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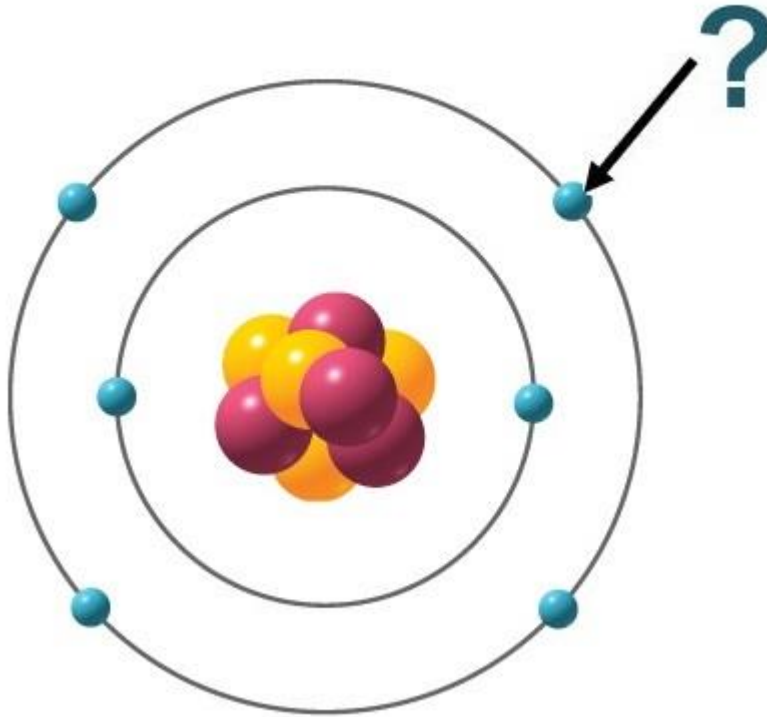
76) Metabolism is the sum of _____ and _____.

- A) inhalation; exhalation
- B) growth; differentiation
- C) anabolism; catabolism
- D) positive; negative feedback
- E) responsiveness; movement

77) Minerals do which of the following?

- A) Contribute to the structure of bones and teeth
- B) Act as fully functional enzymes
- C) Store energy within the body
- D) Act as the monomers of nucleic acids
- E) Form the nuclei of atoms

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78)

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What is indicated by the arrow?

- A) Electron
- B) Proton
- C) Neutron
- D) Anion
- E) Prion

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	Structural formulae	Condensed structural formulae	Molecular formulae
Ethanol	$ \begin{array}{c} \text{H} \quad \text{H} \\ \quad \\ \text{H}-\text{C}-\text{C}-\text{OH} \\ \quad \\ \text{H} \quad \text{H} \end{array} $	$\text{CH}_3\text{CH}_2\text{OH}$	$\text{C}_2\text{H}_6\text{O}$
Ethyl ether	$ \begin{array}{c} \text{H} \quad \quad \text{H} \\ \quad \quad \\ \text{H}-\text{C}-\text{O}-\text{C}-\text{H} \\ \quad \quad \\ \text{H} \quad \quad \text{H} \end{array} $	1	2

79)

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What is the correct condensed structural formula for ethyl ether? (What goes in the box labeled 1?)

- A) CH_3OCH_3
- B) CH_3O
- C) $\text{CH}_3\text{CH}_3\text{OH}$
- D) $\text{C}_2\text{H}_6\text{O}$
- E) $\text{CH}_2\text{CH}_2\text{OH}$

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	Structural formulae	Condensed structural formulae	Molecular formulae
Ethanol	$ \begin{array}{c} \text{H} \quad \text{H} \\ \quad \\ \text{H}-\text{C}-\text{C}-\text{OH} \\ \quad \\ \text{H} \quad \text{H} \end{array} $	$\text{CH}_3\text{CH}_2\text{OH}$	$\text{C}_2\text{H}_6\text{O}$
Ethyl ether	$ \begin{array}{c} \text{H} \quad \quad \text{H} \\ \quad \quad \\ \text{H}-\text{C}-\text{O}-\text{C}-\text{H} \\ \quad \quad \\ \text{H} \quad \quad \text{H} \end{array} $	1	2

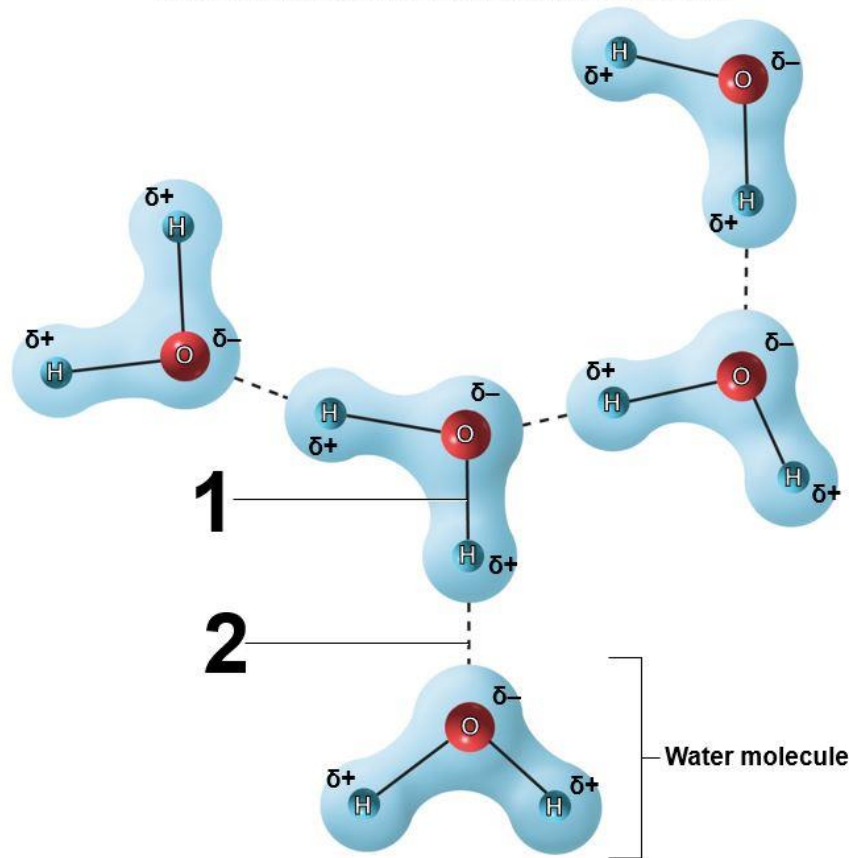
80)

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What is the correct molecular formula for ethyl ether? (What goes in the box labeled 2?)

- A) $\text{C}_2\text{H}_6\text{O}$
- B) $\text{C}_2\text{H}_3\text{O}$
- C) CH_3O
- D) $\text{C}_3\text{H}_6\text{O}$
- E) $\text{C}_3\text{H}_3\text{O}$

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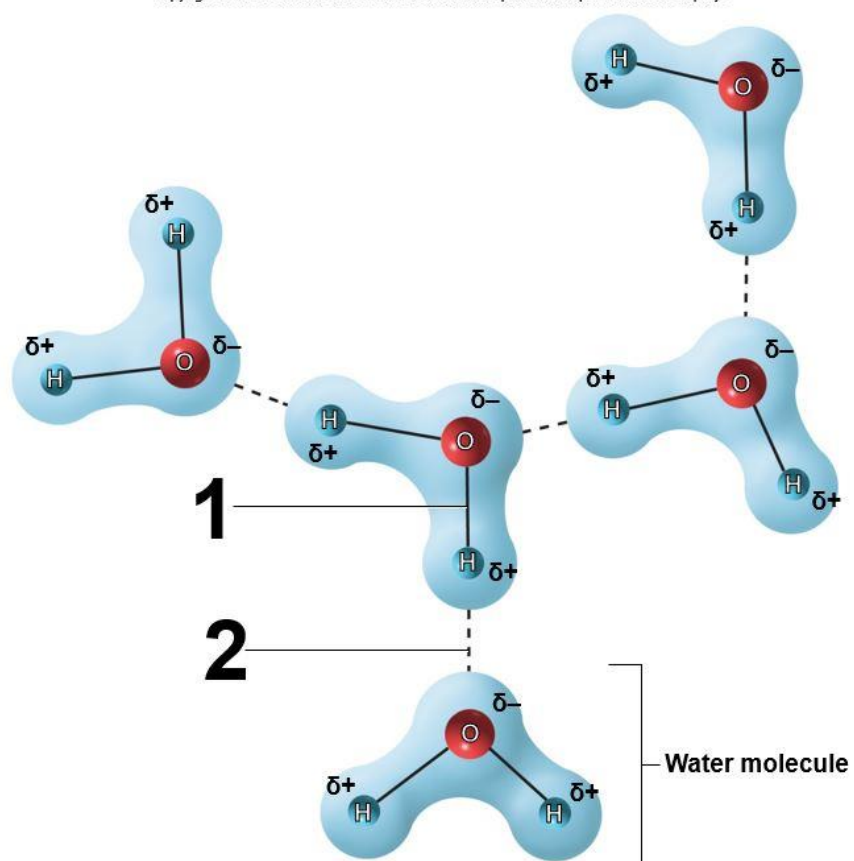


81)

What type of bond is labeled 1?

- A) Covalent
- B) Hydrogen
- C) Ionic
- D) Disulfide
- E) Van der Waals

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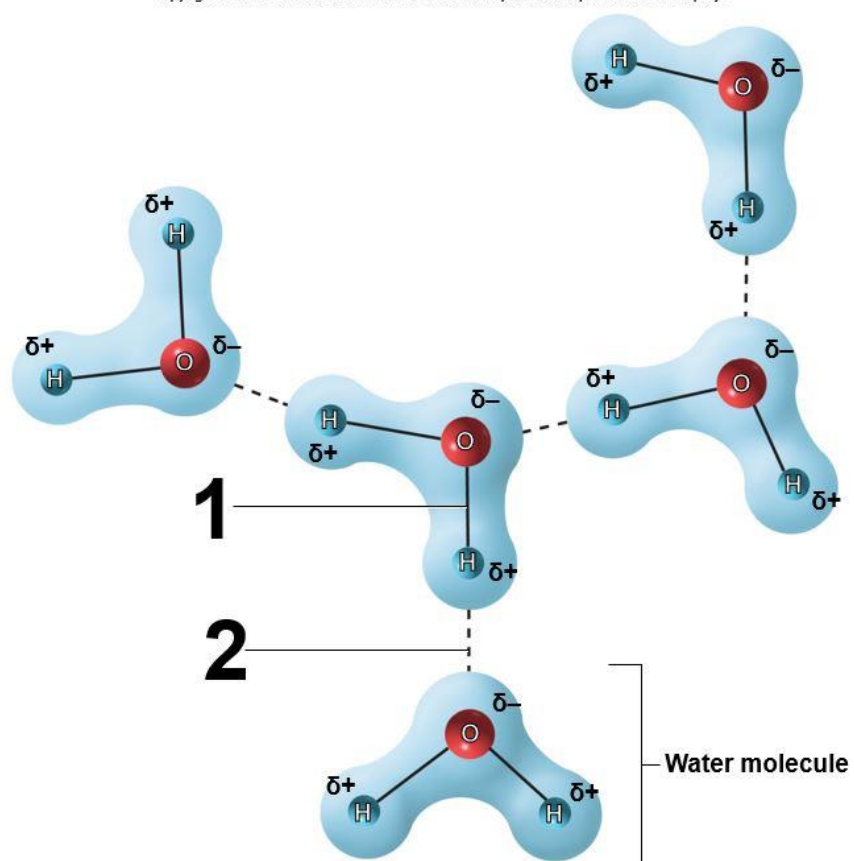


82)

What type of bond is labeled 1?

- A) Single polar covalent bond
- B) Double polar covalent bond
- C) Single nonpolar covalent bond
- D) Double nonpolar covalent bond
- E) Triple covalent bond

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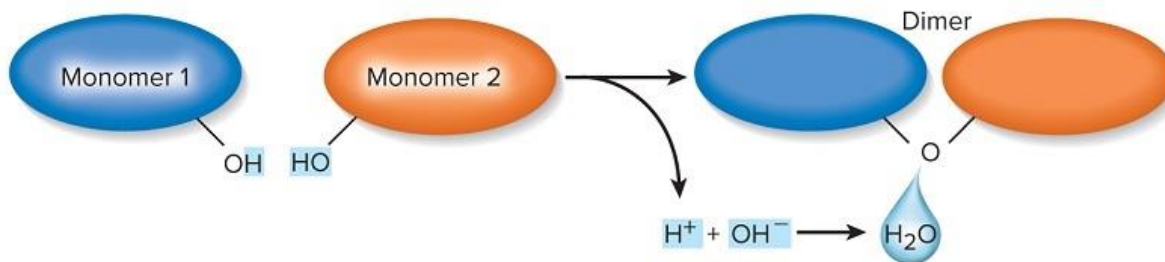


83)

What type of bond is labeled 2?

- A) Hydrogen
- B) Polar covalent
- C) Nonpolar covalent bond
- D) Ionic
- E) Disulfide

84)

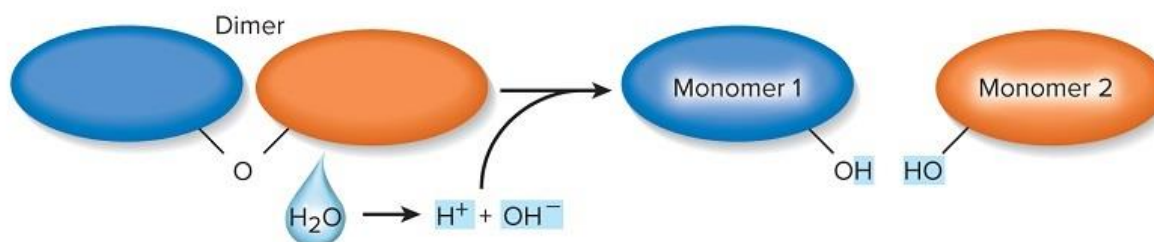


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What type of reaction is shown here?

- A) Dehydration synthesis reaction
- B) Hydrolysis reaction
- C) Exergonic reaction
- D) Catabolic reaction
- E) Oxidation reaction

85)

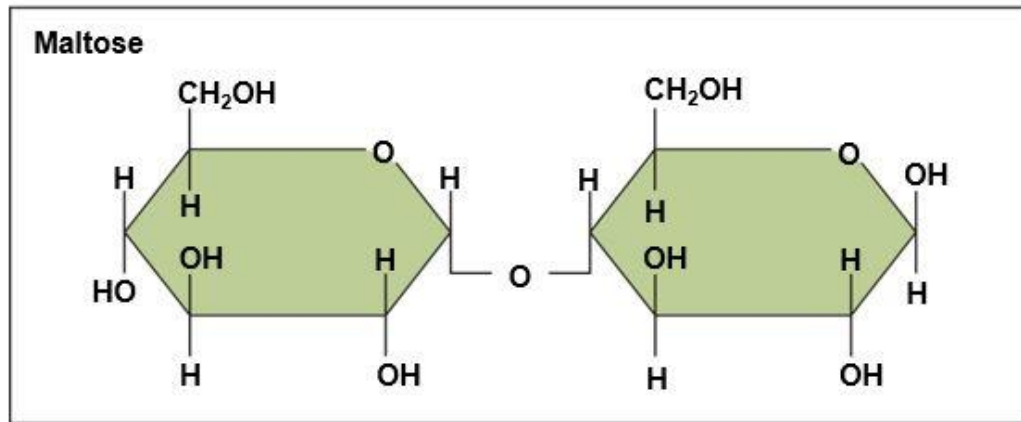


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What type of reaction is shown here?

- A) Hydrolysis reaction
- B) Dehydration synthesis reaction
- C) Endergonic reaction
- D) Anabolic reaction
- E) Reduction reaction

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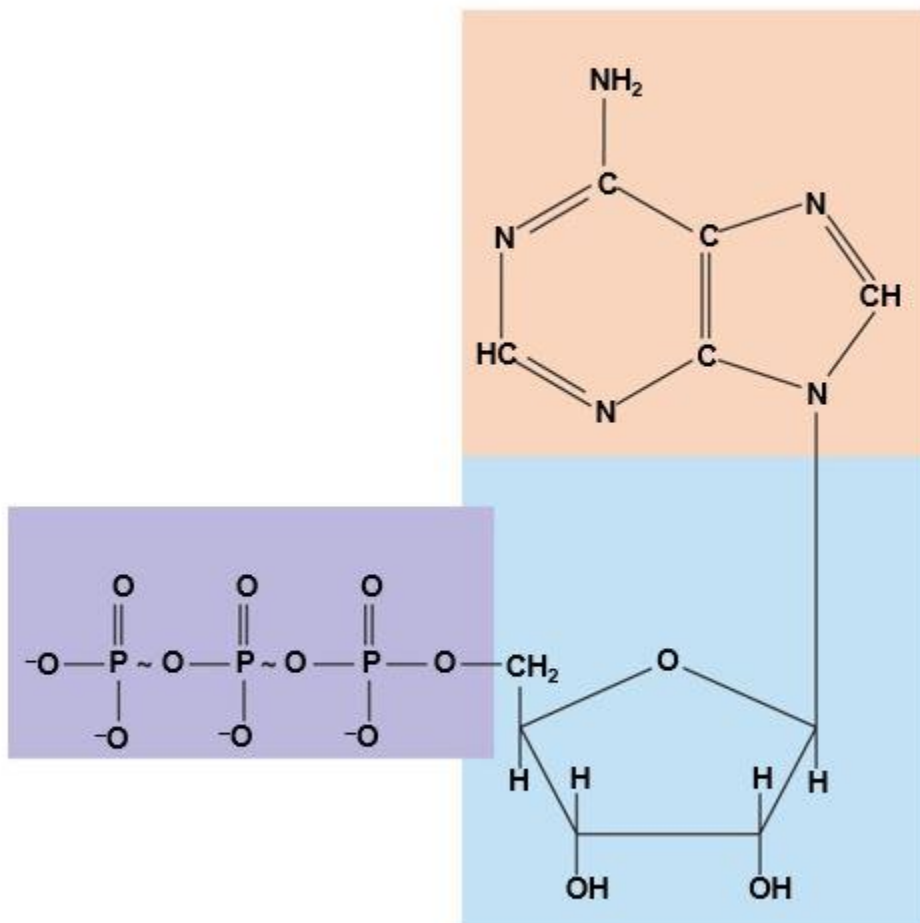
86)

What type of molecule is maltose?

- A) Disaccharide
- B) Monosaccharide
- C) Polysaccharide
- D) Polypeptide
- E) Oligopeptide
- F) Triglyceride

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87)

What molecule is shown here?

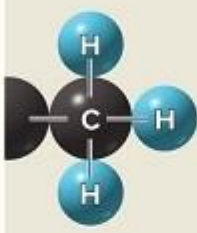
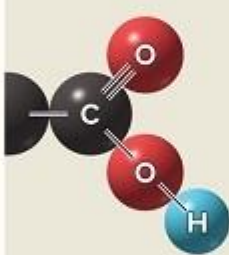
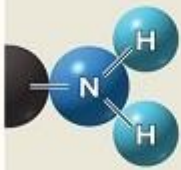
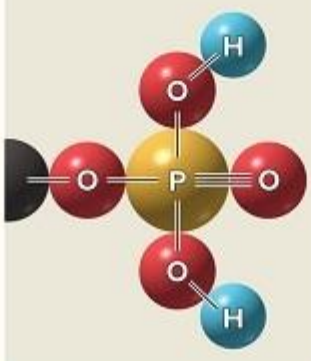
- A) ATP
- B) cAMP
- C) Lecithin
- D) Glucose
- E) Cholesterol

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SECTION BREAK. Answer all the part questions.

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Name and Symbol	Structure	Occurs in
1		Sugars, alcohols
2		Fats, oils, steroids, amino acids
3		Amino acids, sugars, proteins
4		Amino acids, proteins
5		Nucleic acids, ATP

88)

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88.1) Which functional group is labeled 1?

- A) Hydroxyl
- B) Methyl
- C) Carboxyl
- D) Amino
- E) Phosphate

88.2) Which functional group is labeled 2?

- A) Hydroxyl
- B) Methyl
- C) Carboxyl
- D) Amino
- E) Phosphate

88.3) Which functional group is labeled 3?

- A) Hydroxyl
- B) Methyl
- C) Carboxyl
- D) Amino
- E) Phosphate

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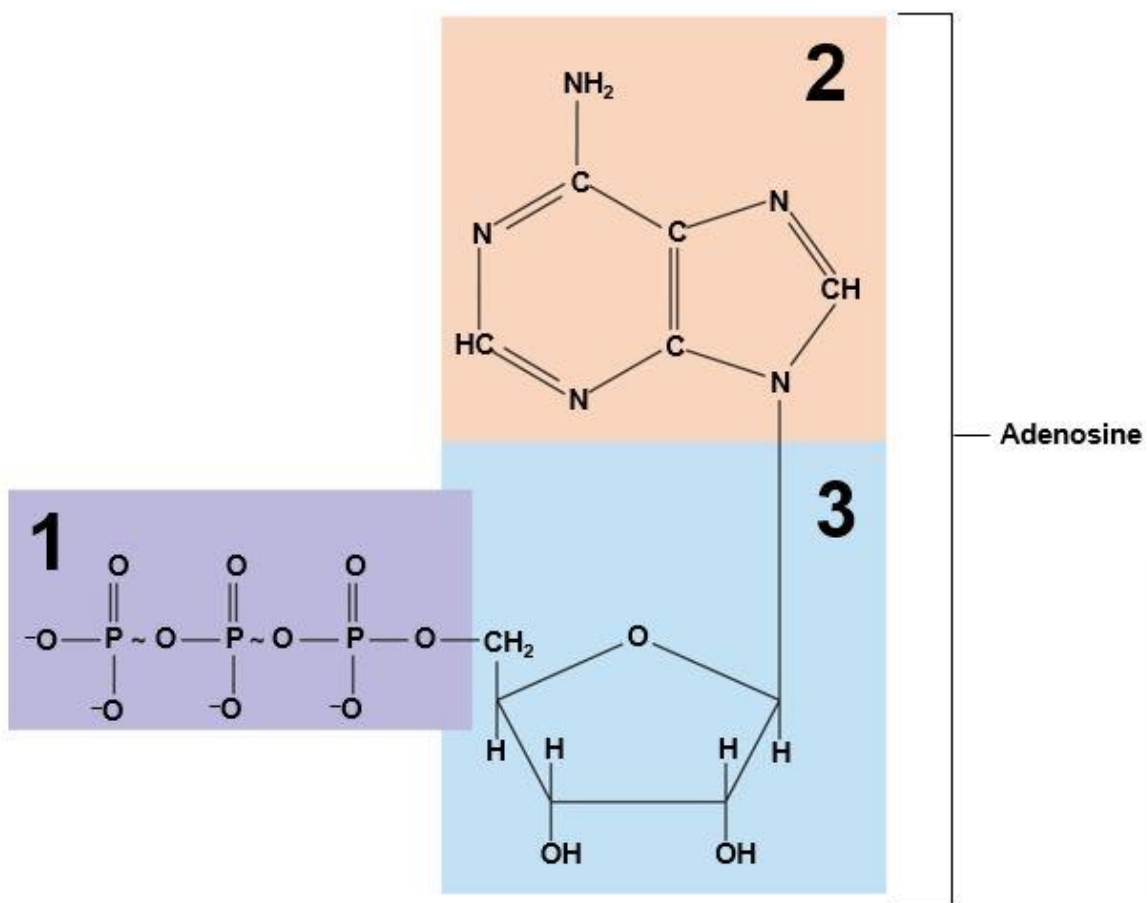
88.4) Which functional group is labeled 4?

- A) Hydroxyl
- B) Methyl
- C) Carboxyl
- D) Amino
- E) Phosphate

88.5) Which functional group is labeled 5?

- A) Hydroxyl
- B) Methyl
- C) Carboxyl
- D) Amino
- E) Phosphate

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(a) Adenosine triphosphate (ATP)

89)

89.1) Identify the structural component of ATP labeled 1.

- A) Triphosphate
- B) Adenine
- C) Ribose
- D) Adenosine
- E) cAMP

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89.2) Identify the structural component of ATP labeled 2.

- A) Triphosphate
- B) Adenine
- C) Ribose
- D) Adenosine
- E) cAMP

89.3) Identify the structural component of ATP labeled 3.

- A) Triphosphate
- B) Adenine
- C) Ribose
- D) Adenosine
- E) cAMP

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Answer Key

Test name: CH-02

- 1) FALSE
- 2) FALSE
- 3) TRUE
- 4) FALSE
- 5) TRUE
- 6) TRUE
- 7) FALSE
- 8) FALSE
- 9) TRUE
- 10) TRUE
- 11) TRUE
- 12) FALSE
- 13) FALSE
- 14) TRUE
- 15) FALSE
- 16) FALSE
- 17) FALSE

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Minerals are inorganic elements extracted from the soil by plants and passed up the food chain to humans.

- 18) [B, C, D]
- 19) D
- 20) A
- 21) B
- 22) E
- 23) C
- 24) B
- 25) D
- 26) A
- 27) C
- 28) E
- 29) C
- 30) B
- 31) E
- 32) E
- 33) A
- 34) D

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- 35) A
- 36) D
- 37) A
- 38) E
- 39) A
- 40) D
- 41) C
- 42) B
- 43) E
- 44) C
- 45) B
- 46) C
- 47) D
- 48) B
- 49) A
- 50) D
- 51) B
- 52) B
- 53) D
- 54) D
- 55) C
- 56) D
- 57) E
- 58) B
- 59) E
- 60) B
- 61) C
- 62) C
- 63) A
- 64) A
- 65) B
- 66) C
- 67) D
- 68) A
- 69) D
- 70) A
- 71) A
- 72) A
- 73) A
- 74) A

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- 75) A
76) C
77) A
78) A
79) A
80) A
81) A
82) A
83) A
84) A
85) A
86) A
87) A
88) Section Break
88.1) A
88.2) B
88.3) C
88.4) D
88.5) E
89) Section Break
89.1) A
89.2) B
89.3) C

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